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Hydrated Crystal Lab Answers

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Hydrated Crystal Lab Answers

In this lab, you will be dehydrating a chemical hydrate and determine the amount of water that will be evaporated away and the anhydrous salt that will be left over. Pre-lab Problem Cobalt (II) Chloride is a hydrated crystal in its solid form. In the lab, you want to determine the formula of this hydrated compound (ie.

Unit 3 Hydrate Lab - JaniceP Per 5 Honchem

The crystals change form, and sometimes color, as the water is driven off. This suggests that water was present as part of the crystal structure.

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Such compounds are called . hydrates. A hydrate that has lost its water is called an . anhydrous salt. For a hydrate, the number of moles of water present per mole of salt is usually some simple, whole number.

Formula of a Hydrate Lab

For the mass of water in hydrated MgSO_4 , just subtract the mass of the anhydrous MgSO_4 from the mass of the hydrated MgSO_4 (the difference must be the water). For moles anhydrous MgSO_4 , divide the...

Hydrated crystals lab help?? | Yahoo Answers

As a good example, copper sulfate is a commonly hydrated crystal. To show that it is hydrated, you would write $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$. By putting a dot in between the two, it indicates that they are waters of...

What is a hydrated crystal? - Answers

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What is a hydrate? Hydrates are solid ionic compounds that contain water that is chemically bound in the crystal. They are crystalline compounds that have a specific number of water molecules trapped within the crystal lattice.

Hydrate Lab by Brandon McDaniel - Prezi

WLHS / Chem / MonsonName Date Per
LAB: Percent Composition of Hydrated Crystals Crystalline compounds that retain water during evaporation are referred to as being hydrated or are said to contain water of hydration. The ratio of moles of water to moles of compound is a small whole number.

LAB: Percent Composition of Hydrated Crystals

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fate hydrate crystals to the crucible. Determine the mass of ... before you leave the lab. Data 'Table Mass of empty crucible and cover 39 Mass of crucible, cover, and magnesium sulfate ... Could the solid be a hydrate? What evidence supports your answer? T b. If, after heating, the solid has a molar mass of 208 g/mol and a formula of ...

Rancho High School

the crystal structure. Crystalline compounds that retain water during

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evaporation are referred to as being hydrated or are said to contain water of hydration. The ratio of moles of water to moles of compound is a small whole number. Example: $\text{BaCl}_2 \cdot 2\text{H}_2\text{O}$ Mole ratio: 1:2 Name: barium chloride dihydrate THIS LAB: $\text{CuSO}_4 \cdot __ \text{H}_2\text{O}$

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Note: The determination and calculation of the formula of a hydrated salt like $\text{MgSO}_4 \cdot 12\text{H}_2\text{O}$ & $128 \cdot 109-113$ odd, pre-lab assignment Wednesday, 8/26 Read chapter 2 complete p. You could buy lead chemistry 11 lab manual

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